

Modeling High-Efficiency Quantum Dot Sensitized Solar Cells

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Inorganic semiconductors constitute promising materials for the development of sensitized solar cells, where an extremely thin semiconductor layer, sandwiched between two wide band gap semiconductors, acts as absorber layer,^{1,2} or in the same configuration than in dye sensitized solar cells (DSCs)³ but using instead of molecular dyes the inorganic semiconductor as light absorbing material. In this last case semiconductor materials are commonly used in the form of quantum dots (QDs)⁴ giving place to quantum dot sensitized solar cells (QDSCs). The use of inorganic semiconductor QDs presents many attractive features for the development of photovoltaic devices. They increase the robustness of the cell and can be prepared in a relatively simple way, with a tunable band gap from the infrared to the ultraviolet region by the control of QD composition and size,⁵ also maintaining beneficial properties of the bulk semiconductors as high molar extinction coefficient⁵ and large intrinsic dipole moments.^{6,7} In addition, the interactions between semiconductor QDs and other compounds can be also advantageous in nanocomposites with enhanced light collecting properties, expanding the light wavelength absorption range^{8,9} and/or reducing the overall recombination in the nanocomposite supracollector.^{10,11}

Although the conversion efficiency, η , of QDSCs currently lags significantly behind maximum efficiencies obtained for DSCs, the gap has been quickly reduced in the past few years and efficiencies close to 4% under full 1 sun illumination have been reported.^{12,13} This fact is significant because the optimal QDSCs configuration has not been obtained yet, and issues as the opti-

ABSTRACT With energy conversion efficiencies in continuous growth, quantum dot sensitized solar cells (QDSCs) are currently under an increasing interest, but there is an absence of a complete model for these devices. Here, we compile the latest developments in this kind of cells in order to attain high efficiency QDSCs, modeling the performance. CdSe QDs have been grown directly on a TiO₂ surface by successive ionic layer adsorption and reaction to ensure high QD loading. ZnS coating and previous growth of CdS were analyzed. Polysulfide electrolyte and Cu₂S counterelectrodes were used to provide higher photocurrents and fill factors, FF. Incident photon-to-current efficiency peaks as high as 82%, under full 1 sun illumination, were obtained, which practically overcomes the photocurrent limitation commonly observed in QDSCs. High power conversion efficiency of up to 3.84% under full 1 sun illumination ($V_{oc} = 0.538$ V, $j_{sc} = 13.9$ mA/cm², FF = 0.51) and the characterization and modeling carried out indicate that recombination has to be overcome for further improvement of QDSC.

KEYWORDS: sensitized solar cells · quantum dots · cadmium selenide · cadmium sulphide · titanium dioxide · SILAR · impedance spectroscopy

cal absorption, charge recombination, hole scavenging, electrolyte, recombination, and series resistance (counterelectrode effect) need to be improved.^{14,15} So far research in this field has focused mainly on showing the validity of the concept and increasing efficiencies and photocurrents of these cells, which were rather low in the beginning. Different semiconductor materials, semiconductor preparation, attaching modes, surface treatments, electrolyte configurations, counter-electrodes, to cite some examples, have been checked in order to improve the performance of QDSCs, but there are few examples of a complete characterization and modeling of photovoltaic devices,^{14,16,17} beyond the conventional current–potential (J – V) curves and incident photon-to-current efficiency (IPCE) measurements. As a consequence no clear model and characterization method have been developed to identify unambiguously the limiting factors in these devices and the real effect in the cell performance of each

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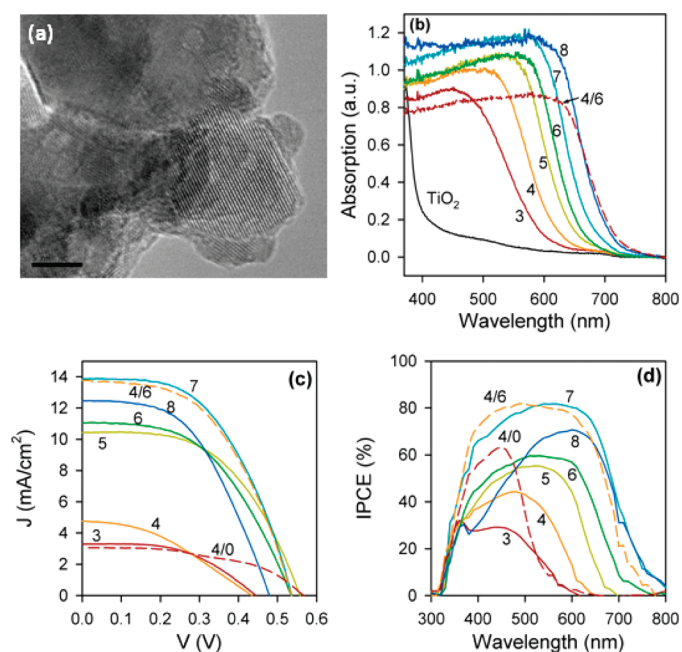


Figure 1. Photovoltaic performance of QDSCs: (a) high-resolution TEM image of CdSe-coated electrode using seven SILAR cycles; (b) absorption spectra of QD sensitized electrodes; (c) J – V curves and (d) IPCE of QDSCs. All the graphics show the behavior depending on the number SILAR cycles. Single number in the legend represents the number of CdSe SILAR cycles. The notation X/Y is referred to CdS/CdSe samples where X and Y are the number of CdS and CdSe SILAR cycles, respectively. The expected j_{sc} for each cell according to the integration of the IPCE curve and the solar irradiance has been compared with the measured j_{sc} , obtaining variations lower than 8%.

of the modifications that have been explored.

Here we show that by compiling the current developments in QDSCs it is possible to overcome the commonly observed photocurrent limitation in these devices. CdSe QDs have been grown directly on a TiO₂ surface by successive ionic layer adsorption and reaction (SILAR)¹⁶ to ensure high QD loading. ZnS coating^{14,18,19} and previous growth of CdS²⁰ were analyzed. Polysulfide electrolyte^{12,20} and Cu₂S counterelectrodes^{18,21} were used to provide higher photocurrents and fill factors, FF. IPCE peaks as high as 82%, under full 1 sun illumination, were obtained, which practically overcomes the photocurrent limitation commonly observed in QDSCs. High power conversion efficiency of up to 3.84% under full 1 sun illumination ($V_{oc} = 0.538$ V, $j_{sc} = 13.9$ mA/cm², FF = 0.51) was obtained. In addition, a careful optoelectronic characterization of the device performance allowed us to model the characteristic features of QDSCs in comparison with the conventional DSCs, showing that there are fundamental differences between the recombination processes of these two cells, as it has been previously suggested¹⁵ and is demonstrated in this work.

RESULTS AND DISCUSSION

In the development of QDSCs two main approximations for the preparation of high harvesting material have been employed: the use of presynthesized colloidal

TABLE 1. Cell Performance^a

number of SILAR cycles CdS/CdSe	V_{oc} (V)	j_{sc} (mA/cm ²)	FF	η (%)	ΔE_{CB-TiO_2} (eV)	β
0/3	0.451	3.3	0.49	0.73	0	0.44
0/4	0.301	5.0	0.42	0.63	0.03	
0/5	0.573	10.3	0.53	3.13	0.03	0.45
0/6	0.550	11.3	0.48	2.98	0.05	0.39
0/7	0.538	13.9	0.51	3.83	0.16	0.32
0/8	0.479	12.6	0.51	3.09	0.16	0.31
4/6	0.548	13.8	0.50	3.78	0.125	0.29
4/0	0.569	3.06	0.48	0.84	0.03	0.27

^aPhotovoltaic properties of the manufactured QDSCs, photocurrent j_{sc} , open circuit voltage V_{oc} , fill factor FF, and efficiency η , as a function of the different surface modifications tested under standard conditions (100 mW/cm² AM 1.5G). The displacement of the TiO₂ CB observed by IS (see Figure 2a,b) and β parameter obtained from β -recombination model^{27,28} are also included.

idal QDs^{17,18} and the direct growth of QDs on a TiO₂ surface by chemical bath deposition (CBD)^{19,20} or SILAR.¹⁶ Colloidal QDs present good performance but low QD loading and CBD higher loading but can present QD aggregation,¹⁴ which introduces a deleterious effect in cell performance.¹⁵ SILAR offers the possibility of high QD loading with low aggregation (see Figure 1a and Supporting Information section S1). To prepare the QDSCs, TiO₂ nanostructured electrodes with two different layers have been employed, one with small TiO₂ particles to ensure high QD loading and another one with bigger particles to increase the light scattering and consequently the light harvesting. After SILAR sensitization, sensitized electrodes present a uniform distribution of QDs along the sample's cross section (see Supporting Information section S2).

The absolute absorption of the sensitized samples increases with the number of SILAR cycles (Figure 1b). In addition the absorption onset shifts to longer wavelengths with the number of SILAR cycles indicating that (i) a reduction of the quantum confinement⁵ occurs and (ii) for the thicker deposits the increased absorption near the bandgap, with smaller absorption coefficient, also contributes to the red shift.¹⁵ QDSCs with different CdSe SILAR cycles were prepared and characterized with J – V curves (see Figure 1c and Table 1). All the samples were coated with ZnS to reduce the recombination of electrons in the TiO₂ to the polysulfide electrolyte^{14,19} with a consequent increase in the photocurrent. On the other hand, Cu₂O was used as counter-electrode material to enhance the charge transfer between counter-electrode and polysulfide,^{18,21} reducing the series resistance of the device and increasing the FF to a value of ~ 0.5 (see Table 1), significantly higher than the reported values of 0.3–0.4 using Pt or Au as counter-electrodes.¹⁸ The combination of these strategies allow us to obtain high QDSCs with a maximum efficiency of $\eta = 3.84\%$ under full 1 sun illumination for seven CdSe SILAR cycles, using mask and no antireflective layer. The IPCE of the prepared samples (Figure 1d) increases with the number or SILAR cycles

and the onset shifts to longer wavelengths, in good correlation with absorption measurements (Figure 1b). This trend is broken for a number of cycles higher than seven, and a reduction of IPCE and cell performance is observed for eight cycles (see Figure 1). For seven cycles an IPCE peak as high as 82% was obtained, which is one of the highest reported IPCE for this kind of cells.

To establish a model of QDSCs performance impedance spectroscopy (IS) measurements were carried out (Figure 2a,b). IS spectra were analyzed using an analogous model as that employed in DSC characterization²² (Figure 2c). It is important to remark that the equivalent circuit presented in Figure 2c reflects the physical processes of carrier transport, charge transfer, and recombination that take place in the device. The equivalent circuit elements have the following physical meaning:²² C_{μ} ($=c_{\mu} \cdot L$, where L is the TiO_2 layer thickness) is the chemical capacitance that stands for the change of electron density as a function of the Fermi level, and it monitors the distribution of traps states in the bandgap of the TiO_2 semiconductor. R_{tr} ($=r_{tr} \cdot L$) is the electron transport resistance in the TiO_2 , directly related to the reciprocal conductivity. R_r ($=r_r \cdot L$) is the recombination resistance, a charge-transfer resistance related to recombination of electrons at the TiO_2 /electrolyte interface. R_r is inversely proportional to both the recombination rate and the density of electrons in TiO_2 . R_s is a series resistance accounting for the transport resistance of the transparent conducting oxide and the connection setup. Z_d is the diffusion impedance of the redox species in the electrolyte (we have not considered this effect in this paper). R_{CE} represents the charge transfer resistance at the counter-electrode/electrolyte interface. C_{CE} is the interfacial capacitance at the counter-electrode/electrolyte interface. The first three mentioned elements (C_{μ} , R_{tr} , and R_r) are denoted in lowercase letters in Figure 2c meaning the element per unit length for a film of thickness L , because they are distributed in a repetitive arrangement of a transmission line. The physical meaning of this network corresponds to the impedance of diffusion and recombination.²³ The IS model presented allows an easy correlation between the impedance spectra and the physical processes, permitting to distinguish separately each effect.

The obtained IS spectra (Figure 2a,b) present the Warburg behavior characteristic of a transmission line²³ (straight line at frequencies higher than the recombination arc), allowing us to obtain transport parameters of electrons in the TiO_2 . The decrease of the recombination resistance, R_r , with the applied voltage can be also visualized as the decrease of the semicircle radius. Using

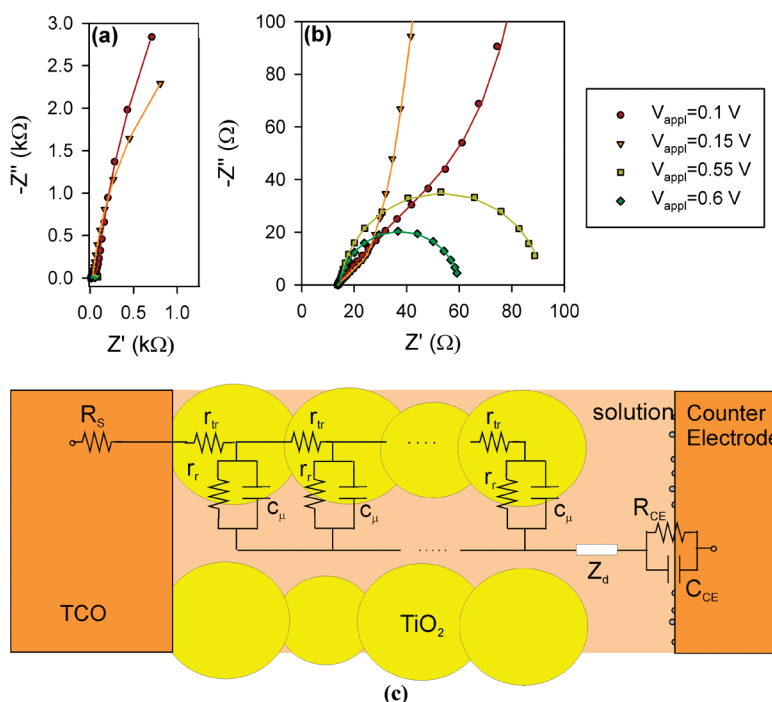


Figure 2. (a) Impedance measurement of QDSCs (symbols) using seven CdSe SILAR cycles and ZnS coating at different applied forward bias, V_{appl} , under dark conditions. (b) Zoom of graph a. Transmission line pattern, observed as a straight line in the high frequency range, is clearly observed for $V_{\text{appl}} = 0.1$ and 0.15 V. Solid lines correspond to the fit using (c) an analogous equivalent circuit as the one employed in DSCs,²² but considering a short circuit in the electrolyte, $Z_d = 0 \Omega$.

polysulfide electrolyte commonly prevents the observation of the Warburg behavior due to the high transfer resistance with the Pt counter-electrode,¹⁴ because the arc due to the counter-electrode at high frequencies is large enough to hide this behavior, but the use of Cu_2S counter-electrode reduces dramatically these charge transfer resistances (Figure 3b). Figure 3b shows the dramatic effect in the charge transfer resistance between the counter-electrode and the polysulfide electrolyte when Cu_2O is used instead of Au or Pt counter-electrodes. Au and Pt counter-electrodes have been checked with electrodes sensitized with CdSe QDs grown by chemical bath deposition (CBD). We have previously shown that the material of the counter-electrode does not affect the parameters of the sensitized electrode,¹⁴ and *vice versa*, so the preparation process of the QD does not affect the counter-electrode transfer resistance. On the other hand, the series resistance (Figure 3a) is independent of the applied voltage as is expected; its values range from 2 to $3.5 \Omega \cdot \text{cm}^2$ for all the samples except for the one prepared with 3 SILAR. These values are comparable with the values obtained for good performing DSCs. In this case of 3 SILAR cycles the higher R_s is produced because this sample was measured without a metallic contact soldered onto the FTO surface.

The observation of the transmission line in the impedance spectra permits a complete characterization of the QDSCs and also allows us to obtain different cell

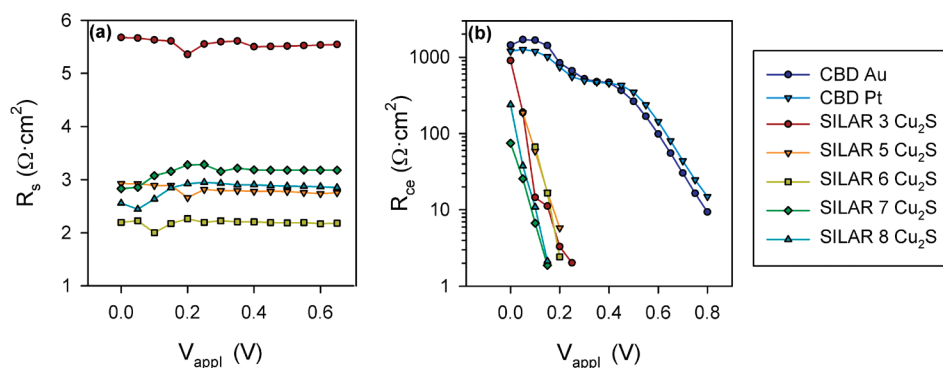


Figure 3. (a) Series resistance, R_s , and (b) counter-electrode resistance, R_{ce} , as function of applied forward bias, V_{appl} , of solar cells with different number of CdSe SILAR cycles and ZnS coating.

parameters as a function of the voltage drop in the sensitized electrode V_F . IS enables extracting the voltage drop in the sensitized electrode, V_F , at each applied potential, V_{appl} , by subtracting the effect of the series resistance and counter-electrode charge transfer resistance on both R_r and C_μ as follows: $V_F = V_{appl} - V_s - V_{ce}$, where V_s and V_{ce} are the potential drop at the series resistance and at the counter-electrode, respectively. V_F is proportional to the rise of the Fermi level of electrons in TiO_2 , $V_F = (E_{Fn} - E_{F0})/q$, where q is the positive elementary charge and E_{Fn} and E_{F0} are the electron Fermi level and the electron Fermi level at the equilibrium, respectively. To analyze the recombination resistance, R_r , on the basis of a similar number of electrons (*i.e.*, the same distance between the electron Fermi level, E_{Fn} , and conduction band (CB) of TiO_2 , E_{CB}) the shift of instead of CB (indicated for each sample in Table 1) has been corrected in Figure 4, where the voltage scale is $V_{F,shifted}$. The criterion for the modified scale is that the chemical capacitances of all the analyzed samples overlap (Figure 4b) because the chemical capacitance²⁴ is directly related with the difference $E_{CB} - E_{Fn}$, by the relation $C_\mu \propto \exp[-(E_{CB} - E_{Fn})/(k_B T)]$. The same shift applied to the chemical capacitance has been applied to the other parameters shown in Figure 4.

The analysis of the chemical capacitance, C_μ , points out to a downward displacement of the TiO_2 conduction band (CB) that in-

creases with the number of CdSe cycles (Figure 4a).

This downward shift has been previously observed after ZnS and/or molecular dipole coating.²⁵ As in the pre-

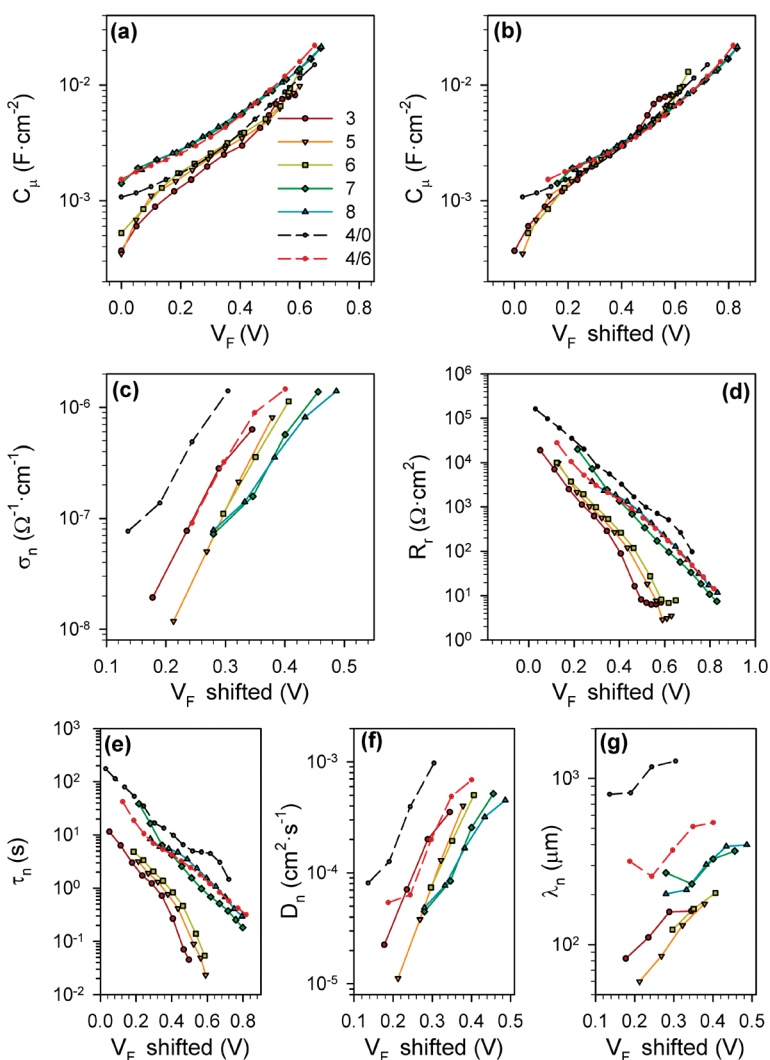


Figure 4. Impedance characterization of QDSCs. Behavior of different parameters of the sensitized electrode obtained by IS, depending on the number SILAR cycles as function of the voltage drop in the electrolyte, V_F or $V_{F,shifted}$ (where the displacement of the TiO_2 CB has been removed). Chemical capacitance, C_μ , as function of (a) V_F and (b) $V_{F,shifted}$. (c) Electron conductivity, σ_n . (d) Recombination resistance R_r . (e) Electron lifetime, τ_n . (f) Electron diffusion coefficient in TiO_2 , D_n . (g) Small perturbation diffusion length in TiO_2 , λ_n .

vious report, now, the increasing amount of CdSe with the number of SILAR cycles prevents the direct contact of the polysulfide electrolyte ($\text{pH} \approx 12$) with TiO_2 , which decreases the effective “observed” pH of the TiO_2 surface, consequently reducing the TiO_2 CB position. The downward displacement of CB can be removed by a voltage shift (Figure 4b) as it has been commented before. With this correction cells can be consequently compared in the same conditions of number of electrons in TiO_2 .

Some authors use the integrated electron density (from the chemical capacitance) to compare the recombination rates, but we believe that using the corrected voltage is a more powerful procedure, as this method allows the simultaneous analysis of the IS and $J-V$ data. A systematic procedure to realize this approach is being developed.²⁶ The number of CdSe SILAR cycles affects the recombination resistance, R_r (Figure 4d), with an increase of R_r with the number of cycles due to a higher loading of CdSe QDs that act themselves as a blocking layer. On the other hand, from the transport resistance R_{tr} it is possible to determine the electron conductivity in the TiO_2 as $\sigma_n = L/R_{tr} \cdot S$, where L is the thickness of TiO_2 and S is the area. σ_n decreases with the number of SILAR cycles (Figure 4c), and this diminution can be attributed to the filling of TiO_2 pores. As the pore size decreases with the number of SILAR cycles, the diffusion into the pores of the chemical species in the electrolyte, allowing a good screening, is reduced, affecting σ_n . Knowing the values of C_{μ} , R_r , and R_{tr} makes it possible to calculate important electron parameters in the device, such as lifetime, τ_n , diffusion coefficient, D_n , and the small perturbation diffusion length, λ_{nr} ,²² see Figure 4 panels e, d, and f, respectively.

The increase of R_r with the amount of QDs produces an increase of lifetime that is partially compensated with a reduction in the D_n as a consequence of the decrease of σ_n . Finally the calculated small perturbation diffusion length values are significantly higher than the TiO_2 thickness ($\sim 14 \mu\text{m}$), indicating that the prepared QDSCs are not transport limited as is expected in a good performing solar cell. As explained elsewhere,²⁷ the small perturbation diffusion length $\lambda_n = (D_n \cdot \tau_n)^{1/2}$ is well-defined, while a general definition of diffusion length is not unambiguous. In the β -recombination model²⁷ the recombination in DSC is proportional to n^β , where n is the electron density in the TiO_2 and β is a parameter. A model based on charge transfer *via* a distribution of surface states on the TiO_2 predicts $\beta = 1/2 + T/T_0$ where T is the temperature and T_0 relates to exponential distribution of traps in TiO_2 ,²⁸ with a minimum value of $\beta = 0.5$. In DSCs β has values that commonly lie between 0.5–0.7. In the β -recombination model the slope of $\ln(R_r)$ vs V_F allows the calculation of β , taking into account the relation:^{27,28} $R_r = R'_0 \exp(-q\beta V_F / (k_B T))$, where q is the positive elemental charge, T is the temperature, k_B is the Boltz-

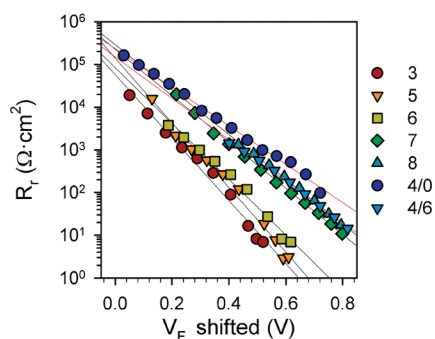


Figure 5. Experimental data (symbols), from Figure 4d, have been fitted to a straight line (solid black line for CdSe and solid red line for CdS/CdSe). A single number in the legend represents the number of CdSe SILAR cycles. The notation X/Y is referred to CdS/CdSe samples where X and Y are the number of CdS and CdSe SILAR cycles, respectively.

mann’s constant, and R'_0 is a pre-exponential factor. In DSCs β can vary with the applied voltage but in good performing cells β is fairly constant in a wide range of V_F values, as occurs in the QDSCs prepared in this work, see Figure 5. We have obtained β in this range. The exponential shape of recombination resistance also warrants the use of a diode model to describe the solar cell $J-V$ curves.

In the case of CdSe QDSCs (Figure 5), for a low number of SILAR cycles the obtained value is close to 0.5 (Table 1), but this value diminishes until close to 0.3, lower than the β minimum, for seven and eight cycles. This fact has important implications because it indicates that there is a clear evolution in the recombination mode between electrodes with low and high QD loading and also that QDs play an active role in the recombination process. As the TiO_2 surface is becoming covered with the number of SILAR cycles the recombination process changes from a preferential recombination with the electrolyte to a preferential recombination with the QDs.

An additional proof of this conclusion can be obtained by V_{oc} decay measurements. To record the V_{oc} decay we have employed two methods in order to increase the voltage of the sample: (i) open circuit voltage decay (OCVD) where the sample was illuminated and the voltage decay was measured after switching off the light, and (ii) applied bias voltage decay (ABVD) where a voltage in the dark is applied and its decay is analyzed in open circuit conditions. Both measurements give similar results for DSCs, if enough short illumination time is employed in OCVD in order to avoid a significant increase of the temperature.²⁹ Taking care of this point we have made both measurements in QDSCs and calculated the lifetime τ_n ²⁹ (Figure 6). There is a good agreement between τ_n obtained by IS and that obtained by ABVD, see inset Figure 6, carried out both under dark conditions. But lower electron life is measured with OCVD. As QDs are the light sensible part of the device, this difference between light and dark measurements indicates that they play an active role in the

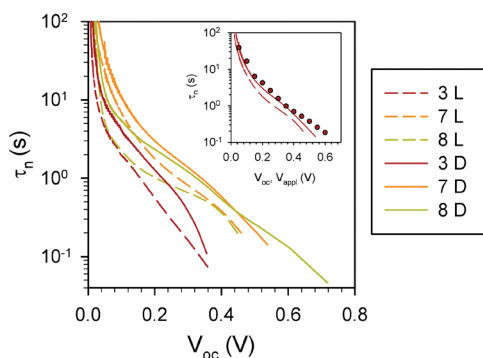


Figure 6. Open circuit voltage decay (OCVD) and applied bias voltage decay (ABVD) characterization of QDSCs. Comparison of the electron lifetime between the photoinduced OCVD (L) and the applied voltage ABVD (D) for QDSCs samples with different number of SILAR cycles. Inset: Comparison between the electron lifetimes extracted from OCVD and IS measurements for the QDSCs prepared with seven SILAR cycles.

recombination process. Illumination produces electron–hole pairs that can be trapped in QD surface states, enhancing the recombination of electrons in the TiO₂ with holes in the QDs or be back injected in electron trap states in the QD, producing a recombination pathway that it is practically nonexistent in good performing DSCs.

If QDs present this active role in recombination, there should be expected also an influence of the QD material. To investigate this we have analyzed the use of CdS as sensitizer. This material is specially interesting because it is extremely easy to prepare and also because it has been reported that its combined use with CdSe can enhance cell performance.^{20,30} An example of the QDSCs using CdS is reported in Figure 1b–d (see also Supporting Information S3). CdS presents lower photocurrent than CdSe cells because its band gap is wider (Figure 1c,d). The combined use of CdS and CdSe allows an increase of IPCE in the short wavelength region (Figure 1d) and efficiency close to the maximum obtained for seven CdSe SILAR cycles is obtained with four and six successive SILAR cycles of CdS and CdSe (Figure 1c,d and Table 1).

From the point of view of recombination it is very interesting to note the large difference between CdS and CdSe sensitized layers. Comparing CdS₄ with CdSe₃, where the number indicates the number of SILAR cycles, only a slight downward shift of the CdS sensitized TiO₂ is observed (0.03 eV, Table 1 and Figure 3a).

MATERIALS AND METHODS

Preparation on Sensitized TiO₂. The electrode configuration was a transparent layer DSL 18NR-T (20 nm average particle size) and a scattering layer WERO-4 (300–400 nm particle size distribution). The FTO (SnO₂:F) coated glass was previously covered by a compact layer of TiO₂ deposited by spray pyrolysis of titanium(IV)bis(acetoacetonato) di(isopropanoxy)ate. These electrodes were sintered at 450 °C for 30 min. The thickness of the

In contrast CdS₄ sample presents R_r and σ_n sensibly higher than CdSe₃. Both variations go toward an increase of cell performance, indicating that CdS allows a better cell performance than CdSe, but the cell is limited by the absorption spectral range of CdS. This absorption range can be expanded using CdSe on top of CdS (see Figure 1d for the case of CdS₄/CdSe₆). In this sample we have obtained R_r and σ_n parameters better than those for CdSe₆ or even CdSe₇ from the point of view of enhanced cell operation. The efficiency obtained for sample CdS₄/CdSe₆ is very close to the one obtained for CdSe₇ (Table 1). A compromise is needed between absorption and pore filling for the combined use of CdS and CdSe, as in the case of CdSe. The use of CdS increases cell performance in the case of six CdSe SILAR cycles if three or four CdS SILAR cycles are employed, but decreases when five cycles are used (Supporting Information section S4). No further improvement of cell performance has been obtained using CdS and seven CdSe SILAR cycles. But, the results reported indicate that in a more open structure than the one used in this study the combined use of CdS and CdSe has a higher potential for good performing cells than CdSe alone provided that the pore filling limitation can be overcome.

A deep analysis of the recombination process in CdS samples points out that cells with only four CdS SILAR cycles present a β lower than 0.3 (Supporting Information section S3), indicating that in these conditions the recombination is already preferential with CdS instead of with the electrolyte. We have demonstrated that the type and size of TiO₂ covering of QDs plays a fundamental role in the QDSCs recombination; further characterization of cyclic voltammetry under dark conditions also points in the same direction (Supporting Information section S5).

CONCLUSIONS

We have shown that it is possible to prepare high performing QDSCs that overcome the photocurrent limit, obtaining QDSCs with a record IPCE. We have given a model for the characterization of QDSCs and demonstrated the role of light absorbing QDs in the recombination process. Further increase in QDSC performance will require an increase of V_{oc} and FF, and to attain this objective a careful control of the recombination through QDs is needed.

mesoporous electrodes was around 14 μm (9 μm of transparent layer and 5 μm of scattering layer) measured by scanning electron microscopy (SEM) with a JSM-7000F JEOL FEG-SEM system, see Supporting Information section S2.

The mesoporous TiO₂ electrodes were *in situ* sensitized by CdS and CdSe QDs grown by SILAR. For CdS growth the electrodes were successively immersed in two different solutions for 1 min each: one consisting of 0.02 M Cd(NO₃)₂ dissolved in methanol, another of 0.02 M Na₂S in methanol/Milli-Q ultrapure

water (1:1). Following each immersion, rinsing and drying was undertaken using pure methanol and a N₂ gun, respectively. All these processes constitute one SILAR cycle. A SILAR of CdSe presents more problems. In order to obtain the Se precursor, we followed the method developed by the group of Grätzel¹⁶, 0.03 M Cd(NO₃)₂ dissolved in ethanol, and another method incorporating the *in situ* generated 0.03 M Se²⁻ in ethanol (see ref 16 for more details). For sensitization electrodes were dipped successively in these solutions inside a glovebox under N₂ atmosphere. Following each immersion, rinsing and drying was undertaken using pure ethanol and a N₂ gun, respectively. In CdS/CdSe samples after CdS sensitization, electrodes were introduced into the glovebox were the CdSe SILAR cycles were carried out. All the samples analyzed in this study have been coated with ZnS, after SILAR sensitization, by twice dipping alternately into 0.1 M Zn(CH₃COO)₂ and 0.1 M Na₂S solutions for 1 min/dip, rinsing with Milli-Q ultrapure water between dips.¹⁹ At least two cells have been prepared at each condition obtaining a variation on cell performance among cells lower than 15%, for cells not presenting problems during the ensemble process (see an example in Supporting Information section S6).

Solar Cell Configuration. The cells were prepared by assembling the counter-electrode and a QD-sensitized electrode using a scotch spacer (thickness 50 μm) and with a droplet (10 μL) of polysulfide electrolyte. Polysulfide electrolyte was 1 M Na₂S, 1 M S, and 0.1 M NaOH solution in Milli-Q ultrapure water. We have used Cu₂S counter-electrodes. The Cu₂S counter-electrodes were prepared by immersing brass in HCl solution at 70 °C for 5 min and subsequently dipping it into polysulfide solution for 10 min, resulting in a porous Cu₂S electrode.²¹ The area of the cells was 0.24 cm².

Electrode and Cell Characterization. The absorption spectra were recorded by a Cary 500 UV–vis Varian spectrometer. For TEM measurements a JEM-2100 electron microscope (JEOL) operated at 200 kV was used. The IPCE measurements were performed employing a 150 W Xe lamp coupled with a monochromator controlled by a computer; the photocurrent was measured using an optical power meter 70310 from Oriol Instruments, using a Si photodiode to calibrate the system. Current–potential curves, cyclic voltammetry, open circuit voltage decay (OCVD) and applied bias voltage decay (ABVD) and impedance spectroscopy measurements were obtained using a FRA-equipped PGSTAT-30 from Autolab. The cells were illuminated using a solar simulator at AM 1.5 G, where the light intensity was adjusted with an NREL-calibrated Si solar cell with a KG-5 filter to 1 sun intensity (100 mW cm⁻²). Impedance spectroscopy measurements were carried out in dark conditions at forward bias, applying a 20 mV AC sinusoidal signal over the constant applied bias with the frequency ranging between 400 kHz and 0.1 Hz. To verify the stability of the cells during IS measurements *J–V* curves were performed before and after IS. No significant variation was observed indicating a good stability during the measurement (see an example in Supporting Information section S7). Measurements were carried out using mask and with no antireflective layer. In OCVD measurements samples were illuminated and the decay of *V*_{oc} was studied as a function of time after the switch-off of the light. In ABVD a current is forced to cross the device under dark, obtaining an applied voltage, then the current is switched off and the voltage decay is measured in open circuit conditions.

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Supporting Information Available: TEM characterization, SEM characterization, characterization of CdS and CdS/CdSe sensitized electrodes, cyclic voltammetry, and cell reproducibility. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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